TEST

"Practice Test (1B-Periods)

10 Questions

10 min

Topics

Atomic Size, Ionization Energy and Electron Affinity, Melting and Boiling Point, Electrical Conductivity + Metallic and non-metallic character, Oxidation States + Hydration Energy, Halides & Oxides

Start Test

09 : 57
1/10
10 min
Hint
Q : For which molecule the bonding pair of electrons is equally shared between the atoms?
A
HF
B
HCl
C
H2
D
BF3
1
2
3
4
5
6
7

09 : 55
2/10
10 min
Hint
Q : Which group of periodic table shown generally abnormal trends of ionization energy.
A IIA and VA
B IIIA and IVA
C IIIA and VIA
D VIIIA
1
2
3
4
5
6
7

09 : 53
3/10
10 min
Hint
Q : Which of the followings are isoelectronic ions?
A O^{-2}, F^{-}
B Na^{+}, F^{-}
C Na^{+}, Al^{+3}
D All of these
1
2
3
4
5
6
7

09 : 51
4/10
10 min
Hint
Q : Which of the following elements forms acidic oxide only
A
Na
B
Al
C
Mg
D
Br
1
2
3
4
5
6
7

09 : 42
5/10
10 min
Hint
Q : Choose the amphoteric oxide
A
Na_2O
B
BaO
C
Al_2O_3
D
SO_3
1
2
3
4
5
6
7

09 : 39
6/10
10 min
Hint
Q : The oxidation number which is not shown by halogens
A -1
B +1
C +2
D +7
1
2
3
4
5
6
7

09 : 37
Q
7/10
10 min
Hint
Q : Which of the following shows minimum electrical conductance
A
Na
B
Al
C
Mg
D
Cu
1
2
3
4
5
6
7

09 : 34
8/10
10 min
Hint
Q : Lanthanide contraction is related to decrease in
A
Atomic number
B
Atomic mass
C
Atomic energy
D
Atomic size
4
5
6
7
8
9
10

09 : 33

9/10 | 10 min | Hint

Q : The electronic configurations of some elements are given below. The element with highest electron affinity is

A $1s^2, 2s^2, 2p^6, 3s^2, 3p^5$

B $1s^2, 2s^2, 2p^6, 3s^2, 3p^1$

C $ls^2, 2s^2, 2p^5$

D $1s^2, 2s^2, 2p^2$

4 5 6 7 8 9 10

09 : 30
10/10
10 min
Hint
Q : The correct atomic size order is
A Li > Mg
B Be > Al
C B > Si
D Ba > Na
4
5
6
7
8
9
10

"Practice Test (
Correct
Unattempted
Incorrect
Q
1/10
Q : For which molecule the bonding pair of electrons is equally shared between the atoms?
A
HF
B
HCl
C
H2
D
BF3
1
2
3
4
5
6
7

"Practice Test (
Correct
Unattempted
Incorrect
Q
2/10
Q : Which group of periodic table shown generally abnormal trends of ionization energy.
A
IIA and VA
B
IIIA and IVA
C
IIIA and VIA
D
VIIIA
1
2
3
4
5
6
7

"Practice Test (
Correct
Unattempted
Incorrect
3/10
Q : Which of the followings are isoelectronic ions?
A O^{-2}, F^{-}
B Na^{+}, F^{-}
C Na^{+}, Al^{+3}
D All of these
1 2 3 4 5 6 7

"Practice Test (
Correct
Unattempted
Incorrect
Q
4/10
Q : Which of the following elements forms acidic oxide only
A
Na
B
Al
C
Mg
D
Br
1
2
3
4
5
6
7

"Practice Test (
Correct
Unattempted
Incorrect
Q
5/10
Q : Choose the amphoteric oxide
A
Na_2O
B
BaO
C
Al_2O_3
D
SO_3
1
2
3
4
5
6
7

"Practice Test (
Correct
Unattempted
Incorrect
Q
6/10
Q : The oxidation number which is not shown by halogens
A -1
B +1
C +2
D +7
1
2
3
4
5
6
7

"Practice Test (
Correct
Unattempted
Incorrect
Q
7/10
Q : Which of the following shows minimum electrical conductance
A
Na
B
Al
C
Mg
D
Cu
1
2
3
4
5
6
7

"Practice Test (
Correct
Unattempted
Incorrect
Q
8/10
Q : Lanthanide contraction is related to decrease in
A
Atomic number
B
Atomic mass
C
Atomic energy
D
Atomic size
4
5
6
7
8
9
10

"Practice Test (

Correct

Unattempted

Incorrect

Q 9/10

Q : The electronic configurations of some elements are given below. The element with highest electron affinity is

A $1s^2, 2s^2, 2p^6, 3s^2, 3p^5$

B $1s^2, 2s^2, 2p^6, 3s^2, 3p^1$

C $ls^2, 2s^2, 2p^5$

D $1s^2, 2s^2, 2p^2$

"Practice Test (
Correct
Unattempted
Incorrect
Q
10/10
Q : The correct atomic size order is
A
Li > Mg
B
Be > Al
C
B > Si
D
Ba > Na
4
5
6
7
8
9
10

TEST

Test Level-1 (1B-Periods)

20 Questions

20 min

Topics

Atomic Size, Ionization Energy and Electron Affinity, Melting and Boiling Point, Electrical Conductivity + Metallic and non-metallic character, Oxidation States + Hydration Energy, Halides & Oxides

Start Test

19 : 56

Q : Select the correct option according to size

A
$Mg^{+} > Mg^{2+}$

B
$Ca^{+} > Ca^{+2}$

C
$Be^{+2} < Be^{+1}$

D All

19 : 54
2/20
20 min
Hint
Q : Which factor does not effect ionization energy across the period
A
Number of shell
B
Nuclear Charge
C
Nature of orbital
D
Both A & C
1
2
3
4
5
6
7

19 : 52
3/20
20 min
Hint
Q :
Mark the correct statement?
A Metallic character increases down the group
B Metallic character increases along a period
C Metallic character decreases down the group
D Metallic character remains the same down the group
1
2
3
4
5
6
7

19 : 49
4/20
20 min
Hint
Q :
Which is the pair of metalloids
A
Na and K
B
F and Cl
C
As and Sb
D
Cu and Au
1
2
3
4
5
6
7

19 : 47
5/20
20 min
Hint
Q : Not a polymeric halide
A
BeCl2
B
CCl4
C
AlCl3
D
GaCl3
1
2
3
4
5
6
7

19 : 45
6/20
20 min
Hint
Q : Which of the following is correct order with respect to covalent nature
A
$PCl_5 > PCl_3$
B
$MgCl_2 > AlCl_3$
C
$PCl_3 > PCl_5$
D
All are correct orders
1
2
3
4
5
6
7

19 : 44

7/20 20 min Hint

Q : For elements of group VA, the oxidation state is equal to

A The number of electrons present in the valance shell

B The number of vacancies available in the valence shell

C The number of all electrons present in all shells

D Both A and B

7

19 : 40

8/20

20 min

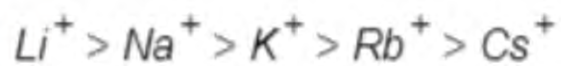
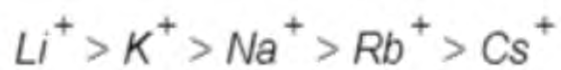
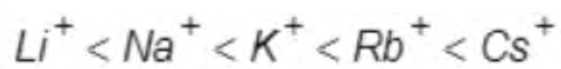
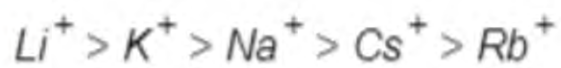
Hint

Q : Which of following is a correct order of degree of hydration in alkali metal ions

A $Li^+ > Na^+ > K^+ > Rb^+ > Cs^+$

B $Li^+ > K^+ > Na^+ > Rb^+ > Cs^+$

C $Li^+ < Na^+ < K^+ < Rb^+ < Cs^+$

D $Li^+ > K^+ > Na^+ > Cs^+ > Rb^+$

6 7 8 9 10 11 12

19 : 38
9/20
20 min
Hint
Q : Keeping the size of atom in view, which order is correct?
A Mg > Sr
B Ba > Mg
C Lu > Ce
D Cl > I
6
7
8
9
10
11
12

19 : 35
10/20
20 min
Hint
Q : which of the following groups show abnormal trends in electron affinity values
A IIA
B VA
C VIIIA
D All of these
6
7
8
9
10
11
12

19 : 34
11/20
20 min
Hint
Q : Which of the following will have zero oxidation state?
A
IA
B
VIA
C
VIIA
D
VIIIA
6
7
8
9
10
11
12

19 : 31
12/20
20 min
Hint
Q : Electrical conductance will be high for
A
Metals
B
Semi-metals
C
None-metals
D
Coinage metals
6
7
8
9
10
11
12

19 : 27
13/20
20 min
Hint
Q : Elements form halides with halogens. Which of these do not form polymeric halide
A
Cs
B
Be
C
Al
D
Ga
10
11
12
13
14
15
16

19 : 25
14/20
20 min
Hint
Q : Which of the following shows minimum electrical conductance?
A Na
B Al
C Mg
D Cu
10
11
12
13
14
15
16

19 : 23
15/20
20 min
Hint
Q : Which ion will have the maximum value of hydration energy
A
Na+
B
Cs+
C
Ba2+
D
Ca2+
10
11
12
13
14
15
16

19 : 21
16/20
20 min
Hint
Q : The elements for which the value of electron affinity is high
A
Gain electrons readily
B
Gain electrons with difficulty
C
Lose electrons less readily
D
Lose electron readily
10
11
12
13
14
15
16

19 : 19
17/20
20 min
Hint
Q : The correct sequence of ionization energy of the elements is
A S < P < Si < Al
B Al < Si < P < S
C Al < Si < S < P
D P < S < Al < Si
11
12
13
14
15
16
17

19 : 16
18/20
20 min
Hint
Q : Which of the following series of elements is listed in order of decreasing atomic radius
A
N, S, F
B
Li, Na, K
C
Na, Mg, Al
D
F, Br, Cl
14
15
16
17
18
19
20

19 : 14
19/20
20 min
Hint
Q :
elting point of I-A and II-A group elements decreases down the group due to __________.
A
Strong electronegativity
B
Strong attractive forces
C
Increment in size
D
High ionization energy
14
15
16
17
18
19
20

19 : 12

20/20

20 min

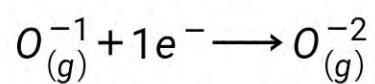
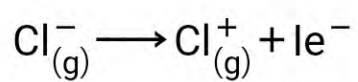
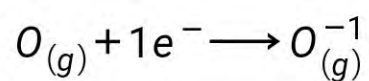
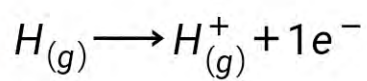
Hint

Q : Which of the following process is exothermic

A $O^{-1}_{(g)} + 1e^{-} \longrightarrow O^{-2}_{(g)}$

B $Cl^{-}_{(g)} \longrightarrow Cl^{+}_{(g)} + le^{-}$

C $O_{(g)} + 1e^{-} \longrightarrow O^{-1}_{(g)}$

D $H_{(g)} \longrightarrow H^{+}_{(g)} + 1e^{-}$

TEST RESULT

Test Level-1 (1B-Periods)

20

20 min

08-Aug-2020

0 sec

Result Detail

Correct	0
Incorrect	0
Unattempted	20

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
1/20
Q : Select the correct option according to size
A
$Mg^{+} > Mg^{2+}$
B
$Ca^{+} > Ca^{+2}$
C
$Be^{+2} < Be^{+1}$
D
All
1
2
3
4
5
6
7

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
2/20
Q : Which factor does not effect ionization energy across the period
A
Number of shell
B
Nuclear Charge
C
Nature of orbital
D
Both A & C
1
2
3
4
5
6
7

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
3/20
Q :
Mark the correct statement?
A
Metallic character increases down the group
B
Metallic character increases along a period
C
Metallic character decreases down the group
D
Metallic character remains the same down the group
1
2
3
4
5
6
7

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
4/20
Q :
Which is the pair of metalloids
A
Na and K
B
F and Cl
C
As and Sb
D
Cu and Au
1
2
3
4
5
6
7

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
5/20
Q : Not a polymeric halide
A
BeCl2
B
CCl4
C
AlCl3
D
GaCl3
1
2
3
4
5
6
7

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
6/20
Q : Which of the following is correct order with respect to covalent nature
A
$PCl_5 > PCl_3$
B
$MgCl_2 > AlCl_3$
C
$PCl_3 > PCl_5$
D
All are correct orders
1
2
3
4
5
6
7

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
7/20
Q : For elements of group VA, the oxidation state is equal to
A
The number of electrons present in the valance shell
B
The number of vacancies available in the valence shell
C
The number of all electrons present in all shells
D
Both A and B

Q : Which of following is a correct order of degree of hydration in alkali metal ions

A $Li^{+} > Na^{+} > K^{+} > Rb^{+} > Cs^{+}$

B $Li^{+} > K^{+} > Na^{+} > Rb^{+} > Cs^{+}$

C $Li^{+} < Na^{+} < K^{+} < Rb^{+} < Cs^{+}$

D $Li^{+} > K^{+} > Na^{+} > Cs^{+} > Rb^{+}$

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
9/20
Q : Keeping the size of atom in view, which order is correct?
A
Mg > Sr
B
Ba > Mg
C
Lu > Ce
D
Cl > I
5
6
7
8
9
10
11

Test Level-1 (1B-
Correct
Unattempted
Incorrect
10/20
Q : which of the following groups show abnormal trends in electron affinity values
A IIA
B VA
C VIIIA
D All of these
5
6
7
8
9
10
11

Test Level-1 (1B-
Correct
Unattempted
Incorrect
11/20
Q : Which of the following will have zero oxidation state?
A IA
B VIA
C VIIA
D VIIIA
11 12 13 14 15 16 17

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
12/20
Q : Electrical conductance will be high for
A
Metals
B
Semi-metals
C
None-metals
D
Coinage metals
11
12
13
14
15
16
17

Test Level-1 (1B-
Correct
Unattempted
Incorrect
13/20
Q : Elements form halides with halogens. Which of these do not form polymeric halide
A Cs
B Be
C Al
D Ga
11
12
13
14
15
16
17

Test Level-1 (1B-
Correct
Unattempted
Incorrect
14/20
Q : Which of the following shows minimum electrical conductance?
A Na
B Al
C Mg
D Cu
11
12
13
14
15
16
17

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
15/20
Q : Which ion will have the maximum value of hydration energy
A
Na+
B
Cs+
C
Ba2+
D
Ca2+
11
12
13
14
15
16
17

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
16/20
Q : The elements for which the value of electron affinity is high
A
Gain electrons readily
B
Gain electrons with difficulty
C
Lose electrons less readily
D
Lose electron readily
11
12
13
14
15
16
17

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q 17/20
Q : The correct sequence of ionization energy of the elements is
A S < P < Si < Al
B Al < Si < P < S
C Al < Si < S < P
D P < S < Al < Si
11 12 13 14 15 16 17

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
18/20
Q : Which of the following series of elements is listed in order of decreasing atomic radius
A
N, S, F
B
Li, Na, K
C
Na, Mg, Al
D
F, Br, Cl
14
15
16
17
18
19
20

Test Level-1 (1B-
Correct
Unattempted
Incorrect
Q
19/20
Q :
elting point of I-A and II-A group elements decreases down the group due to ___________.
A
Strong electronegativity
B
Strong attractive forces
C
Increment in size
D
High ionization energy
14
15
16
17
18
19
20

Test Level-1 (1B-

Correct

Unattempted

Incorrect

Q 20/20

Q : Which of the following process is exothermic

A $O^{-1}_{(g)} + 1e^{-} \longrightarrow O^{-2}_{(g)}$

B $Cl^{-}_{(g)} \longrightarrow Cl^{+}_{(g)} + 1e^{-}$

C $O_{(g)} + 1e^{-} \longrightarrow O^{-1}_{(g)}$

D $H_{(g)} \longrightarrow H^{+}_{(g)} + 1e^{-}$

TEST

Test Level-2 (Topic 1B)

30 Questions

25 min

Topics

Atomic Size, Ionization Energy and Electron Affinity, Melting and Boiling Point, Electrical Conductivity + Metallic and non-metallic character, Oxidation States + Hydration Energy, Halides & Oxides

Start Test

24 : 57
1/30
25 min
Hint
Q : Which of the following has minimum hydration energy
A Li+
B Ca++
C Na+
D K+
1
2
3
4
5
6
7

24 : 55
2/30
25 min
Hint
Q : Which of the following oxide is classified incorrectly
A Na_2O—Neutral
B CO_2—Acidic
C NO_2—Acidic
D BeO—Amphoteric
1
2
3
4
5
6
7

24 : 53

3/30 25 min Hint

Q : Which of the following electronic configuration represents atoms of element having the highest 2^{nd} Ionization energy

A $1s^2 2s^2 2p^4$

B $1s^2 2s^2 2p^5$

C $1s^2 2s^2 2p^6 3s^1$

D $1s^2 2s^2 2p^6 3s^2$

24 : 51
4/30
25 min
Hint
Q : Which of the following is most acidic
A
Mn_2O_7
B
Mn_2O_3
C
MnO_2
D
MnO
1
2
3
4
5
6
7

24 : 49
5/30
25 min
Hint
Q : Which element has highest melting point
A
Be
B
C
C
B
D
N
1
2
3
4
5
6
7

24 : 48
6/30
25 min
Hint
Q : The p-block element among the following which is smallest in size is
A
Li
B
P
C
Na
D
Cl
1
2
3
4
5
6
7

24 : 46
7/30
25 min
Hint
Q : Which of the following does not form polymeric halide
A Cs
B Be
C Al
D Ga
1
2
3
4
5
6
7

24 : 43

8/30 25 min Hint

Q : The increasing order of the first ionization enthalpies of the elements P, S and Cl is

A. Cl < S < P

B. P < S < Cl

C. P < Cl < S

D. S < P < Cl

5 6 7 8 9 10 11

24 : 41
Q 9/30
25 min
Hint
Q : The most ionic chloride is
A NaCl
B CsCl
C $BaCl_2$
D $MgCl_2$
5
6
7
8
9
10
11

24 : 39
10/30
25 min
Hint
Q : The element whose normal oxide is amphoteric in nature
A
B
B
Al
C
Mg
D
C
5
6
7
8
9
10
11

24 : 37

11/30 25 min Hint

Q : The melting and boiling point along 2nd and 3rd period

A Increases from left to right

B Decrease from left to right

C 1st increases upto group IVA and then decreases

D 1st decreases upto group IV A and then increases

24 : 34
12/30
25 min
Hint
Q : The ionization energy of boron is less than that of beryllium because
A
Atomic size of B > Be
B
Shielding effect of B < Be
C
Atomic size of B < Be
D
Removal of e⁻ from p-subshell
9
10
11
12
13
14
15

24 : 32

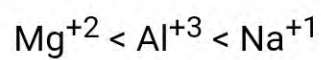
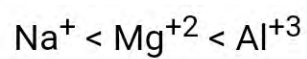
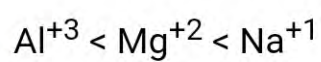
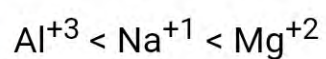
13/30 25 min Hint

Q : Which among the following is the correct order of increasing ionic radius

A $Al^{+3} < Na^{+1} < Mg^{+2}$

B $Al^{+3} < Mg^{+2} < Na^{+1}$

C $Na^{+} < Mg^{+2} < Al^{+3}$

D $Mg^{+2} < Al^{+3} < Na^{+1}$

24 : 30
14/30
25 min
Hint
Q : Which of the following shows minimum electrical conductance
A
Na
B
Al
C
Mg
D
Cu
9
10
11
12
13
14
15

24 : 28

15/30 25 min Hint

Q : The ionization energy of nitrogen is more than that of oxygen because of

A More attraction of electrons by the nucleus

B Higher nuclear charge

C The extra stability of half-filled p – orbitals

D The size of nitrogen atom is smaller

13 14 15

24 : 25

16/30 25 min Hint

Q :

An atom has electronic configuration: $1s^2$, $2s^2$, $2p^6$, $3s^2$, $3p^4$ You will place it in

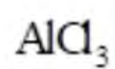
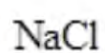
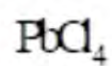
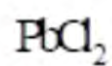
A Group II A , Period 2

B Group VI A , Period 2

C Group II A , Period 3

D Group VI A , Period 3

24 : 24
17/30
25 min
Hint
Q : One of the following factors that has no effect on the 1st ionization energy along the 2nd and 3rd periods
A
Shielding effect
B
Nuclear charge
C
Half filled orbitals
D
Completely filled orbitals
12
13
14
15
16
17

24 : 21
18/30
25 min
Hint
Q : Which of the following is more covalent in nature
A
PbCl2
B
PbCl4
C
NaCl
D
AlCl3
3
14
15
16
17
18
19
20

24 : 19
19/30
25 min
Hint
Q :
Nitrogen oxygen and fluorine have low melting points in 2nd period because
A
They are small size elements
B
They are non-metals
C
They exist as individual molecules
D
They are non-conductors
3
14
15
16
17
18
19
20

24 : 16

20/30 25 min Hint

Q : Generally electronegativity of element in periods _______ and in groups _______

A Increases, Increases

B Increases, Decreases

C Decreases, Decreases

D Decreases, Increases

16 17 18 19 20 21 22

24 : 14
21/30
25 min
Hint
Q : Mg forms Mg+2 but does not form Mg+3 because of
A
Very high 2nd I.E
B
Very high 3rd I.E
C
Very low 2nd I.E
D
Very low 3rd I.E

24 : 13

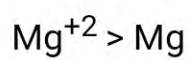
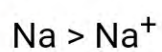
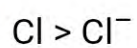
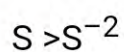
22/30 25 min Hint

Q : The correct order of size is

A $S > S^{-2}$

B $Cl > Cl^{-}$

C $Na > Na^{+}$

D $Mg^{+2} > Mg$

16 17 18 19 20 21 22

24 : 11
23/30
25 min
Hint
Q : The ionic nature of 2nd period chlorides
A
Decreases
B
Increases
C
Remains constant
D
First increases and then decreases
7
18
19
20
21
22
23
24

Q : Among the following outermost electronic configuration of the least electronegative element in a given period is

A) $ns^2 np^5$

B) $ns^2 np^2$

C) $ns^2 np^4$

D) $ns^2 np^3$

24 : 07

25/30 25 min Hint

Q : In which of the following pairs, the ionization energy of first specie is less than that of second

A B, Be

B N, O

C N, P

D P, S

25

24 : 05

26/30

25 min

Hint

Q : The ionic radii (A^0) of N^{3-} , O^{2-} and F^- are respectively

A 1.36,1.40 and 1.71

B 1.36, 1.71 and 1.40

C 1.71, 1.40 and 1.36

D 1.71, 1.36 and 1.40

24 : 03
27/30
25 min
Hint
Q : Which of the following oxides is unlikely to dissolve in aqueous sodium hydroxide
A SiO_2
B MgO
C Al_2O_3
D P_4O_{10}
21
22
23
24
25
26
27

Q : Heat of hydration of gaseous hydrogen ion when it dissolved in water to make infinitely dilute solution is

A $-1075 kJmole^{-1}$

B $-1280 kJmole^{-1}$

C $-1175 kJmole^{-1}$

D $-965 kJmole^{-1}$

23 : 58

29/30 25 min Hint

Q :

Melting and boiling points for first three elements of 3^{rd} period are in order of $Na < Mg < Al$ It is due to

A Increase of charge on metal ions

B Increase in numbers of delocalized electrons

C Increase in the strength of the metallic bonding

D All of these

Q : Sum of number of all elements belonging to 2^{nd} and 3^{rd} periods collectively are

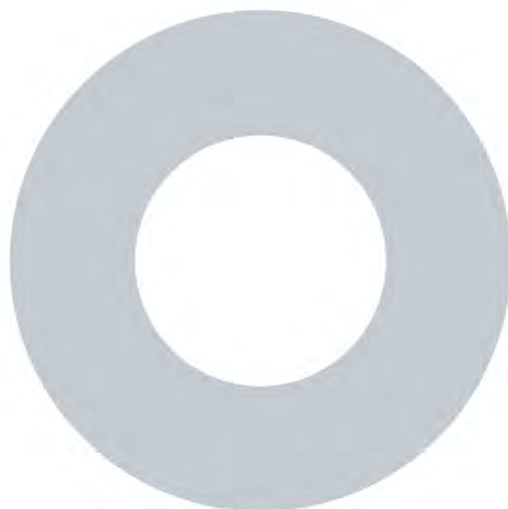
A 10

B 16

C 26

D 36

TEST RESULT

Test Level-2 (Topic 1B)
30
25 min
08-Aug-2020
0 sec
Result Detail

Correct 0
Incorrect 0
Unattempted 30

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
1/30
Q : Which of the following has minimum hydration energy
A
Li^{+}
B
Ca^{++}
C
Na^{+}
D
K^{+}
1
2
3
4
5
6
7

Test Level-2 (To
Correct
Unattempted
Incorrect
2/30
Q : Which of the following oxide is classified incorrectly
A
Na,O—Neutral
B
CO2—Acidic
C
NO2—Acidic
D
BeO—Amphoteric
1
2
3
4
5
6
7

Test Level-2 (To
Correct
Unattempted
Incorrect
Q 3/30
Q : Which of the following electronic configuration represents atoms of element having the highest 2^{nd} Ionization energy
A $1s^2\ 2s^2\ 2p^4$
B $1s^2\ 2s^2\ 2p^5$
C $1s^2\ 2s^2\ 2p^6\ 3s^1$
D $1s^2\ 2s^2\ 2p^6\ 3s^2$

1 2 3 4 5 6 7

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
4/30
Q : Which of the following is most acidic
A
Mn_2O_7
B
Mn_2O_3
C
MnO_2
D
MnO
1
2
3
4
5
6
7

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
5/30
Q : Which element has highest melting point
A
Be
B
C
C
B
D
N
1
2
3
4
5
6
7

Test Level-2 (To
Correct
Unattempted
Incorrect
6/30
Q : The p-block element among the following which is smallest in size is
A Li
B P
C Na
D Cl
1
2
3
4
5
6
7

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
7/30
Q : Which of the following does not form polymeric halide
A
Cs
B
Be
C
Al
D
Ga
1
2
3
4
5
6
7

Test Level-2 (To
Correct
Unattempted
Incorrect
8/30
Q : The increasing order of the first ionization enthalpies of the elements P, S and Cl is
A Cl < S < P
B P < S < Cl
C P < Cl < S
D S < P < Cl
8 9 10 11 12 13 14

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
9/30
Q : The most ionic chloride is
A
NaCl
B
CsCl
C
$BaCl_2$
D
$MgCl_2$
8
9
10
11
12
13
14

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
10/30
Q : The element whose normal oxide is amphoteric in nature
A
B
B
Al
C
Mg
D
C
8
9
10
11
12
13
14

Q : The melting and boiling point along 2^{nd} and 3^{rd} period

A. Increases from left to right

B. Decrease from left to right

C. 1^{st} increases upto group IVA and then decreases

D. 1^{st} decreases upto group IV A and then increases

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
12/30
Q : The ionization energy of boron is less than that of beryllium because
A Atomic size of B > Be
B Shielding effect of B < Be
C Atomic size of B < Be
D Removal of e⁻ from p-subshell
8 9 10 11 12 13 14

Correct

Unattempted

Incorrect

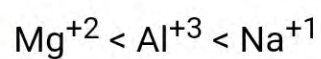
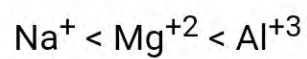
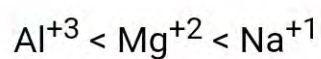
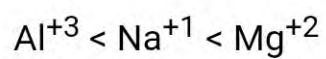
Q 13/30

Q : Which among the following is the correct order of increasing ionic radius

A $Al^{+3} < Na^{+1} < Mg^{+2}$

B $Al^{+3} < Mg^{+2} < Na^{+1}$

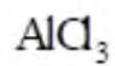
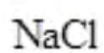
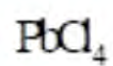
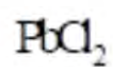
C $Na^{+} < Mg^{+2} < Al^{+3}$

D $Mg^{+2} < Al^{+3} < Na^{+1}$

Test Level-2 (To
Correct
Unattempted
Incorrect
14/30
Q : Which of the following shows minimum electrical conductance
A Na
B Al
C Mg
D Cu
8 9 10 11 12 13 14

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
15/30
Q : The ionization energy of nitrogen is more than that of oxygen because of
A
More attraction of electrons by the nucleus
B
Higher nuclear charge
C
The extra stability of half-filled p – orbitals
D
The size of nitrogen atom is smaller
14
15
16
17
18
19
20

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
16/30
Q :
An atom has electronic configuration: $1s^2, 2s^2, 2p^6, 3s^2, 3p^4$ You will place it in
A Group II A , Period 2
B Group VI A , Period 2
C Group II A , Period 3
D Group VI A , Period 3
14 15 16 17 18 19 20

Test Level-2 (To
Correct
Unattempted
Incorrect
17/30
Q : One of the following factors that has no effect on the 1st ionization energy along the 2nd and 3rd periods
A Shielding effect
B Nuclear charge
C Half filled orbitals
D Completely filled orbitals
14 15 16 17 18 19 20

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
18/30
Q : Which of the following is more covalent in nature
A
$PbCl_2$
B
$PbCl_4$
C
NaCl
D
$AlCl_3$
14
15
16
17
18
19
20

Test Level-2 (To
Correct
Unattempted
Incorrect
19/30
Q :
Nitrogen oxygen and fluorine have low melting points in 2nd period because
A They are small size elements
B They are non-metals
C They exist as individual molecules
D They are non-conductors
14 15 16 17 18 19 20

Test Level-2 (To
Correct
Unattempted
Incorrect
20/30
Q : Generally electronegativity of element in periods _______ and in groups _______
A Increases, Increases
B Increases, Decreases
C Decreases, Decreases
D Decreases, Increases
14
15
16
17
18
19
20

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
21/30
Q : Mg forms Mg+2 but does not form Mg+3 because of
A
Very high 2nd I.E
B
Very high 3rd I.E
C
Very low 2nd I.E
D
Very low 3rd I.E
20
21
22
23
24
25
26

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
22/30
Q : The correct order of size is
A
$S > S^{-2}$
B
$Cl > Cl^{-}$
C
$Na > Na^{+}$
D
$Mg^{+2} > Mg$
20
21
22
23
24
25
26

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
23/30
Q : The ionic nature of 2nd period chlorides
A
Decreases
B
Increases
C
Remains constant
D
First increases and then decreases
20
21
22
23
24
25
26

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
24/30
Q : Among the following outermost electronic configuration of the least electronegative element in a given period is
A
ns2 np5
B
ns2 np2
C
ns2 np4
D
ns2 np3
20
21
22
23
24
25
26

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
25/30
Q : In which of the following pairs, the ionization energy of first specie is less than that of second
A
B, Be
B
N, O
C
N, P
D
P, S
20
21
22
23
24
25
26

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
26/30
Q : The ionic radii (A^0) of N^{3-} , O^{2-} and F^- are respectively
A
1.36,1.40 and 1.71
B
1.36, 1.71 and 1.40
C
1.71, 1.40 and 1.36
D
1.71, 1.36 and 1.40
20
21
22
23
24
25
26

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
27/30
Q : Which of the following oxides is unlikely to dissolve in aqueous sodium hydroxide
A
SiO_2
B
MgO
C
Al_2O_3
D
P_4O_{10}
24
25
26
27
28
29
30

Test Level-2 (To
Correct
Unattempted
Incorrect
28/30
Q : Heat of hydration of gaseous hydrogen ion when it dissolved in water to make infinitely dilute solution is
A $-1075 kJmole^{-1}$
B $-1280 kJmole^{-1}$
C $-1175 kJmole^{-1}$
D $-965 kJmole^{-1}$
24
25
26
27
28
29
30

Q :

Melting and boiling points for first three elements of 3^{rd} period are in order of Na < Mg < Al It is due to

A Increase of charge on metal ions

B Increase in numbers of delocalized electrons

C Increase in the strength of the metallic bonding

D All of these

Test Level-2 (To
Correct
Unattempted
Incorrect
Q
30/30
Q : Sum of number of all elements belonging to 2nd and 3rd periods collectively are
A
10
B
16
C
26
D
36
24
25
26
27
28
29
30